

Write your name here

Surname

Other names

**Pearson**  
**Edexcel GCE**

Centre Number

Candidate Number

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# **Chemistry**

## **Advanced Subsidiary**

### **Paper 1: Core Inorganic and Physical Chemistry**

Friday 27 May 2016 – Morning  
**Time: 1 hour 30 minutes**

Paper Reference  
**8CH0/01**

**You must have:**

Data Booklet  
Calculator

Total Marks

#### **Instructions**

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
  - *there may be more space than you need.*

#### **Information**

- The total mark for this paper is 80.
- The marks for **each** question are shown in brackets
  - *use this as a guide as to how much time to spend on each question.*
- You may use a scientific calculator.
- For questions marked with an **asterisk (\*)**, marks will be awarded for your ability to structure your answer logically showing the points that you make are related or follow on from each other where appropriate.

#### **Advice**

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.
- Show all your working in calculations and include units where appropriate.

**Turn over** ►

**P49837A**

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**PEARSON**

**Answer ALL questions.**

**Some questions must be answered with a cross in a box .**  
**If you change your mind about an answer, put a line through the box  and then mark your new answer with a cross .**

**1** This question is about Group 7 elements and their compounds.

(a) (i) Give the physical states of chlorine and iodine at room temperature and pressure.

(1)

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(ii) Predict the physical state of astatine under these conditions. Justify your answer.

(1)

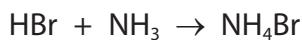
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(b) Write the equation for the reaction of chlorine with cold, dilute sodium hydroxide solution to form bleach. Name this type of reaction.

(2)

Type of reaction .....

(c) Hydrogen bromide gas reacts with ammonia gas



What would be observed during this reaction?

(1)

- A** bubbles
- B** decolorisation
- C** steamy fumes
- D** white smoke



(d) State what is meant by the term electronegativity and hence explain the polarity, if any, of the **bonds** in chlorine trifluoride,  $\text{ClF}_3$ .

(3)

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(e) What is the number of ions in 9.53 g of magnesium chloride,  $\text{MgCl}_2$ ?

[Avogadro constant =  $6.02 \times 10^{23} \text{ mol}^{-1}$ ]

(1)

- A  $6.02 \times 10^{22}$
- B  $1.20 \times 10^{23}$
- C  $1.81 \times 10^{23}$
- D  $6.02 \times 10^{23}$

(Total for Question 1 = 9 marks)



2 This question is about the structure of the atom and isotopes.

The following excerpt is taken from the book *Inorganic Chemistry* by Bailey and Snellgrove, fourth impression 1938.

*"Some of the electrons are also contained in the nucleus, whilst the remainder are ..... arranged in rings revolving round the nucleus ..... The two isotopes [of chlorine] have therefore 18 and 20 electrons respectively in the nucleus and 17 [electrons] external to it."*

(a) Identify and correct **two** errors in the excerpt.

(2)

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.....

(b) What is the structure of a 1+ ion of the carbon-13 isotope?

(1)

- A six protons, six neutrons and five electrons
- B six protons, seven neutrons and six electrons
- C six protons, seven neutrons and five electrons
- D seven protons, six neutrons and six electrons

(c) (i) State what is meant by the term **relative atomic mass**.

(2)

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(ii) A 5.000 g sample of lithium, containing the two isotopes lithium-6 and lithium-7, was found to contain 0.460 g of the isotope lithium-6.

Calculate the relative atomic mass of lithium for this sample. Give your answer to an appropriate number of significant figures.

Isotope	Relative isotopic mass
Lithium-6	6.015
Lithium-7	7.016

(3)

(d) A mass spectrometer was used to analyse a sample of bromine,  $\text{Br}_2$ , with only the  $^{79}\text{Br}$  and  $^{81}\text{Br}$  isotopes present.

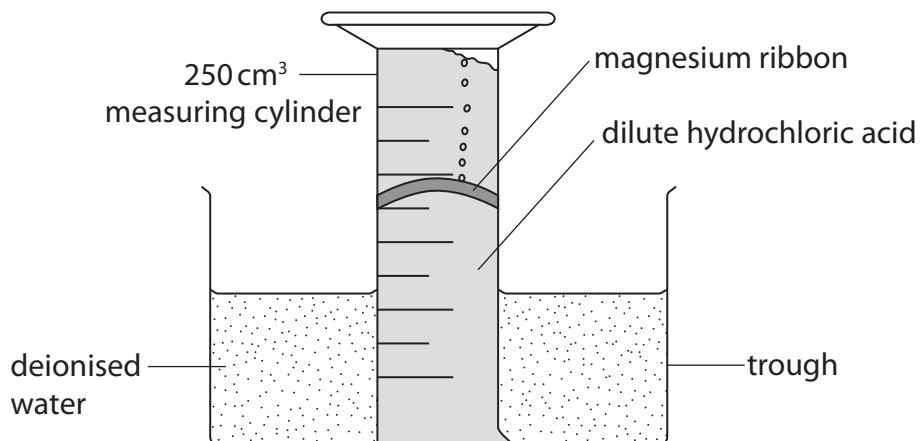
Explain why a very small peak occurs at  $m/z = 80$ .

(2)

(Total for Question 2 = 10 marks)



3 A student used the apparatus in the diagram to determine the molar volume of a gas.



Assume clamps are in place to secure the apparatus

The student used a piece of magnesium ribbon, which was about 5 cm in length, and the dilute hydrochloric acid was in excess. The experiment was repeated three times at 24 °C and the following results were obtained.

	Experiment 1	Experiment 2	Experiment 3
Mass of magnesium / g	0.04	0.04	0.04
Volume of hydrogen gas / cm <sup>3</sup>	31	25	32

The equation for the reaction is



(a) (i) Calculate the number of moles of magnesium used by the student in each experiment.

(1)

(ii) Use your answer from part (a)(i) to deduce the number of moles of hydrogen gas that should be produced.

(1)



(iii) Calculate, using the Ideal Gas Equation, the volume of hydrogen gas, in **cm<sup>3</sup>**, that should be produced in each of these experiments.

$$[pV = nRT \quad R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1} \quad p = 101\,000 \text{ Pa}]$$

(4)

(b) Give a reason why the student repeated the experiment three times.

(1)

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(c) Give three reasons for the difference between your calculated value in (a)(iii) and the actual volumes of hydrogen gas obtained by the student.

For each reason, identify a change to either the apparatus or the chemicals that could be made by the student to improve the result.

(6)

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**(Total for Question 3 = 13 marks)**



P 4 9 8 3 7 A 0 7 2 0

4 The labels on four colourless solutions had fallen off in storage. It was known that the solutions were:

hydrochloric acid  
potassium chloride

magnesium sulfate  
sodium carbonate

In order to identify each solution, a number of tests were carried out.

(a) Solutions can be sprayed into a flame to produce a flame colour identical to that seen in the more conventional method with a solid on a nichrome wire.

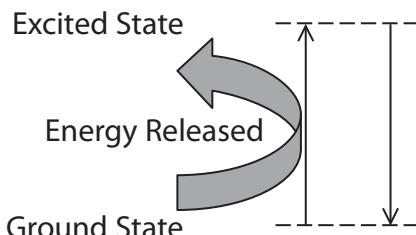
Which solution would produce a lilac flame?

(1)

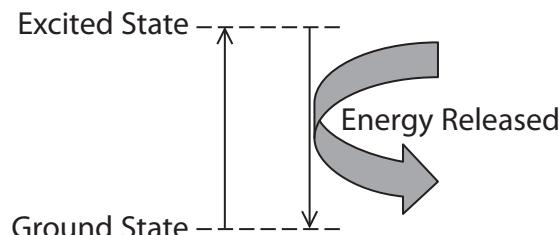
- A hydrochloric acid
- B magnesium sulfate
- C potassium chloride
- D sodium carbonate

(b) Which of the following diagrams best illustrates the electronic transitions that take place during a flame test?

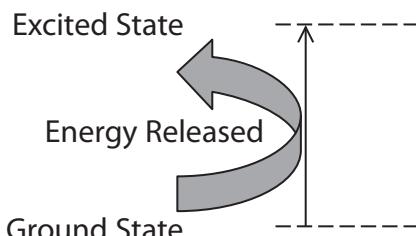
(1)



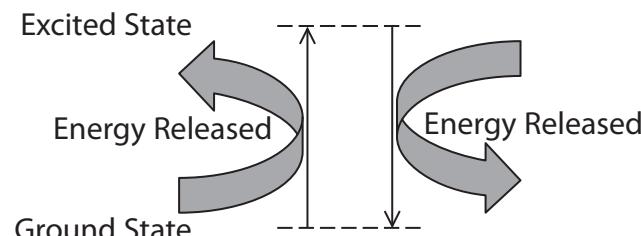
A



B



C



D



(c) Which solution produces a white precipitate with acidified barium chloride solution? (1)

- A hydrochloric acid
- B magnesium sulfate
- C potassium chloride
- D sodium carbonate

(d) Two of the solutions produce the same result on the addition of dilute nitric acid followed by silver nitrate solution.

State the observation with this test and the **two** solutions that give this result.

(2)

Observation .....

Solutions .....

(e) The hydrochloric acid and the sodium carbonate solution react together. State an observation you would make and write the **ionic** equation for the reaction. State symbols are not required.

(2)

Observation .....

Ionic equation

**(Total for Question 4 = 7 marks)**



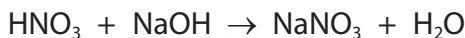
5 A solution of nitric acid,  $\text{HNO}_3$ , of concentration  $100 \text{ g dm}^{-3}$ , can be used to artificially age wood.

A sample of nitric acid, thought to be suitable for this use, was diluted by pipetting  $10.00 \text{ cm}^3$  of this acid into a  $250 \text{ cm}^3$  volumetric flask, adding deionised water and making the solution up to the mark. The solution was thoroughly mixed.

A titration was carried out using this diluted solution of nitric acid. The burette was filled with  $0.0800 \text{ mol dm}^{-3}$  sodium hydroxide solution and  $25.00 \text{ cm}^3$  of the diluted nitric acid was pipetted into each of three conical flasks. The following results were obtained.

	<b>Titration 1</b>	<b>Titration 2</b>	<b>Titration 3</b>
Final burette reading / $\text{cm}^3$	20.50	40.40	20.00
Initial burette reading / $\text{cm}^3$	0.00	20.50	0.00
Volume added / $\text{cm}^3$	20.50	19.90	20.00

The equation for the reaction is



(a) Select the appropriate titres and calculate the mean titre in  $\text{cm}^3$ .

(1)

(b) Calculate the concentration of the **undiluted** nitric acid in  $\text{g dm}^{-3}$ . Give your answer to one decimal place.

Deduce whether this nitric acid is suitable for use in artificially ageing wood.

(5)

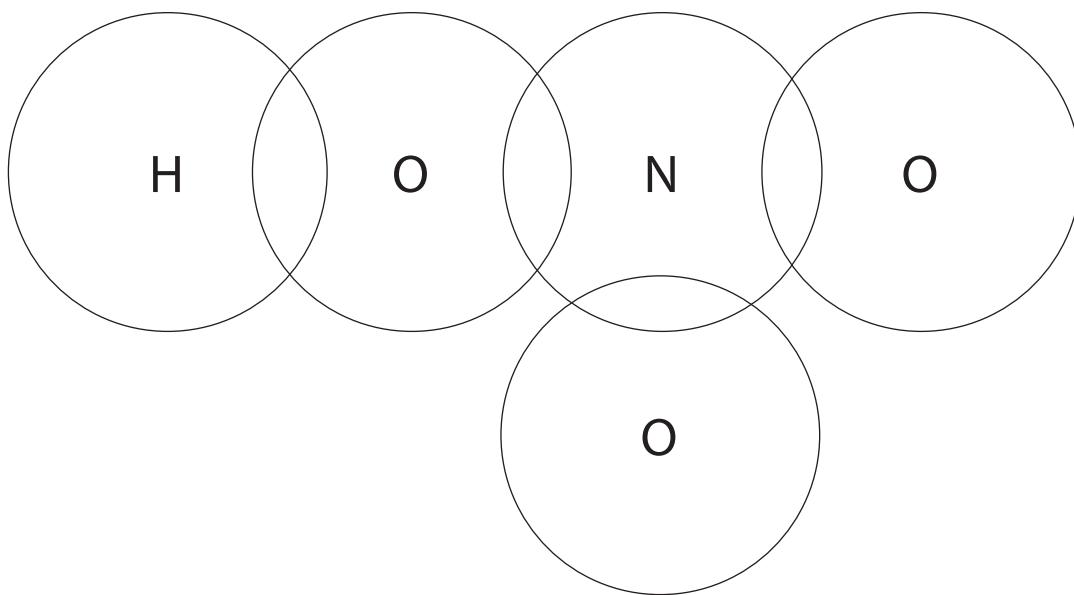


(c) Complete the dot-and-cross diagram for the bonding in nitric acid, showing only outer shell electrons.

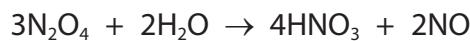
Use (•) for the oxygen electrons,

(x) for the nitrogen electrons and (\*) for the hydrogen electron.

(3)



(d) One possible method for the formation of nitric acid involves the reaction between dinitrogen tetroxide and water.



Calculate the atom economy for the formation of nitric acid from this reaction.

(1)

**(Total for Question 5 = 10 marks)**



6 (a) The diagram shows bond angles in ammonia and water.



Explain why the bond angle in water is less than the bond angle in ammonia.

(2)

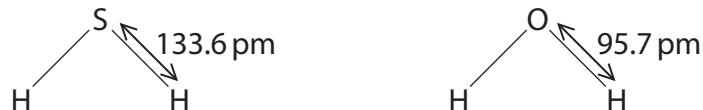
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(b) Explain why the O—H and S—H bond lengths are different.



(3)

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**(Total for Question 6 = 5 marks)**



7 \* (a) A student suggested that the difference in the rates of reaction of strontium and barium with water is due to the difference in the sum of their first and second ionisation energies. Discuss this suggestion.

(6)

(b) Which is the equation for the **third** ionisation energy of barium?

(1)

- A**  $\text{Ba(s)} \rightarrow \text{Ba}^{3+}(\text{g}) + 3\text{e}^-$
- B**  $\text{Ba}^{2+}(\text{s}) \rightarrow \text{Ba}^{3+}(\text{g}) + \text{e}^-$
- C**  $\text{Ba}^{2+}(\text{g}) \rightarrow \text{Ba}^{3+}(\text{g}) + \text{e}^-$
- D**  $\text{Ba}^{3+}(\text{g}) \rightarrow \text{Ba}^{4+}(\text{g}) + \text{e}^-$



(c) Explain why magnesium nitrate,  $\text{Mg}(\text{NO}_3)_2$  decomposes more readily on heating than potassium nitrate,  $\text{KNO}_3$ .

(4)

(d) Some metal carbonates also undergo thermal decomposition.

(i) Draw a diagram of the apparatus that could be used to compare the ease of thermal decomposition of lithium carbonate,  $\text{Li}_2\text{CO}_3$ , and magnesium carbonate,  $\text{MgCO}_3$ .

(2)



(ii) State **one** way in which you would ensure a fair test.

(1)

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.....

(iii) State how data obtained in this experiment could be used to make a comparison.

(1)

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**(Total for Question 7 = 15 marks)**



P 4 9 8 3 7 A 0 1 5 2 0

8 The properties of elements and their compounds are determined by their structure and bonding.

(a) Magnesium oxide has a very high melting temperature.

Which of the following is the best description of its structure and bonding?

(1)

A giant ionic

B giant metallic

C giant covalent

D simple covalent

(b) Sulfur reacts with fluorine to form a number of different compounds.

(i) One compound contains 45.79% sulfur and 54.21% fluorine by mass.

Calculate the empirical formula of this compound.

(2)

(ii) In a dry container, a fluoride of silver reacts with sulfur to produce disulfur difluoride. Complete the equation for this reaction.

State symbols are not required.

(1)



(iii) Explain, by using the oxidation numbers of **all** the atoms, whether or not this is a redox reaction.

(3)



(c) Element **X** has the typical appearance of a metal.

Predict **two** other distinct physical properties that element **X** would exhibit if it is a metal. Explain your choices in terms of structure and bonding.

(4)

**(Total for Question 8 = 11 marks)**

**TOTAL FOR PAPER = 80 MARKS**



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# The Periodic Table of Elements

1 2

3 4 5 6 7 0 (8)  
(18)

1.0  
H  
hydrogen  
1

## Key

relative atomic mass
atomic symbol
atomic (proton) number

6.9 Li lithium 3	9.0 Be beryllium 4	(1)	(2)	10.8 B boron 5	12.0 C carbon 6	14.0 N nitrogen 7	16.0 O oxygen 8	19.0 F fluorine 9	20.2 Ne neon 10
23.0 Na sodium 11	24.3 Mg magnesium 12	(3)	(4)	45.0 Ca calcium 20	47.9 Sc scandium 21	52.0 Ti titanium 22	54.9 Cr chromium 23	55.8 Mn manganese 24	58.7 Fe iron 25
39.1 K potassium 19	40.1 Ca calcium 20	(5)	(6)	91.2 Sr strontium 38	92.9 Y yttrium 39	95.9 Nb niobium 40	95.9 Mo molybdenum 41	[98] Tc technetium 42	101.1 Ru ruthenium 43
85.5 Rb rubidium 37	87.6 Sr strontium 38	(7)	(8)	178.5 Ba barium 56	180.9 Hf hafnium 57	183.8 Ta tantalum 73	186.2 Re rhenium 75	190.2 Os osmium 76	192.2 Ir iridium 77
132.9 Cs caesium 55	137.3 Ba barium 56	(9)	(10)	138.9 La* lanthanum 57	140.9 Hf hafnium 72	141.9 Ta tantalum 73	143.9 W tungsten 74	145.9 Os osmium 75	147.9 Ir iridium 77
[223] Fr francium 87	[226] Ra radium 88	(11)	(12)	[227] Rf rutherfordium 104	[226] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[268] Mt meitnerium 108	[271] Ds darmstadtium 109
		(13)	(14)	[147] Ce cerium 58	144 Pr praseodymium 59	150 Nd neodymium 60	152 Sm promethium 61	157 Eu europium 63	159 Gd gadolinium 64
		(15)	(16)	[242] Am americium 95	[243] Pu plutonium 94	[245] Cm curium 96	[247] Bk berkelium 97	[251] Cf californium 98	[254] Es einsteinium 99
		(17)							[256] Fm fermium 100
									[257] Md mendelevium 101
									[254] No nobelium 102
									[257] Lr lawrencium 103

Elements with atomic numbers 112-116 have been reported  
but not fully authenticated

140 Ce cerium 58	141 Pr praseodymium 59	144 Nd neodymium 60	147 Pm promethium 61	150 Sm samarium 62	152 Eu europium 63	157 Gd gadolinium 64	159 Tb terbium 65	163 Dy dysprosium 66	165 Ho holmium 67	167 Er erbium 68	169 Tm thulium 69	173 Yb ytterbium 70	175 Lu lutetium 71
232 Th thorium 90	[231] Pa protactinium 91	238 U uranium 92	[237] Np neptunium 93	[242] Pu plutonium 94	[243] Am americium 95	[247] Cm curium 96	[245] Bk berkelium 97	[251] Cf californium 98	[254] Es einsteinium 99	[253] Fm fermium 100	[256] Md mendelevium 101	[254] No nobelium 102	[257] Lr lawrencium 103

\* Lanthanide series

\* Actinide series

